Name: _____

Date: _____

Pd: _____

Elements, Compounds & Mixtures Lab

Station 1: Marshmallow Compounds:

Use toothpicks and marshmallows to create compounds listed below. Follow the directions/example in the box.

Marshmallow Key:

Hydrogen (H) – pink

Oxygen (O) – blue

Sodium (Na) – green

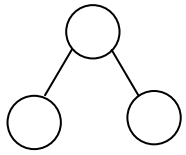
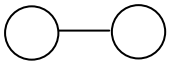
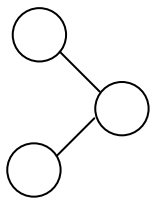
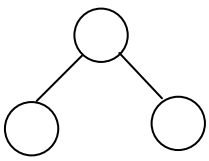
Calcium (Ca) – white

Chlorine (Cl) – yellow

- What do the toothpicks that hold together the atoms represent?

  (Hint: think about the definition of a compound)

_____

Compound List the name of the atoms and the number of each	Gumdrop Model Make the marshmallow compound and color the diagram
H_2O H = Hydrogen (x2) O = Oxygen (x1)	
NaCl	
Na_2O	
$CaCl_2$	

Station 2: Copper

Is This An Element, Mixture or Compound?

This is a(n) _____ because _____

_____

Station 3: Lucky Charms

Is This An Element, Mixture or Compound?

This is a(n) _____ because _____

_____

Station 4: Salt (NaCl - Sodium Chloride)

Is This An Element, Mixture or Compound?

This is a(n) _____ because _____

_____

Station 5: Iron Filings & Pepper

What Kind of Mixture Is This? (Suspension, Colloid, or Solution)

This is a _____ because _____

_____

Try to separate it into parts!

What physical property did you use to separate this mixture? _____

Station 6: Mayonaise on a Plate

What Kind of Mixture Is This? (Suspension, Colloid, or Solution)

This is a _____ because _____

_____

Is this Homogeneous or Heterogeneous? _____

Station 7: Mio Flavored Water

What Kind of Mixture Is This? (Suspension, Colloid, or Solution)

This is a _____ because _____

_____

Is this Homogeneous or Heterogeneous? _____

Station 8: Sugar and Water Solution

In this solution, which material is the SOLUTE? (is being dissolved) _____

Which material is the SOLVENT? (is doing the dissolving) _____

Station 9: Gobstopper In A Dish of Water

What process is happening to the gobstopper? (physical or chemical change) It is a _____

because this type of change results in _____

You'll notice the dye from the candy is mixing with the water creating a solution.

What is the SOLUTE? _____

What is the SOLVENT? _____

Metric Practice:

Complete the metric conversions below: (Remember: What unit are you starting with? Where do you want to go? Count and move your decimal the same number of spaces in the same direction)

Kilo k	Hecto h	Deka da	Grams(g) Meters (m) Liters (L)	deci d	centi c	milli m

3893.2 kg = _____dag

0.0975 hm = _____ cm

4833 L = _____dL

69.856 mg = _____ kg

Lab Reflection:

DESCRIBE three things you learned this week that were demonstrated in this lab.

1. _____

_____

2. _____

_____

3. _____

_____

What are two things that may be confusing/easy to get wrong to some people? (Make sure you explain WHY)

1. _____

_____

2. _____

_____

What did you think of this lab?

_____

_____

_____

_____